

CHEM 204 – General Chemistry Online Lab Recitation

Sections 003

Spring 2024

Instructor: Sam Johnson

E-mail Address: johnsonsj@winthrop.edu

Office: 109F Sims

Office Hours: M, T 4:00 pm - 6:30 PM, or by appointment

Lecture: Tuesdays 6:30-7:20 pm online in Blackboard

Course Credit Hours: 2 credit hour for lab and recitation

Required Textbook:

Textbook: *Chemistry: An Atoms-Focused Approach* by Gilbert, Kirss and Foster, 3rd edition or you may use the free online textbook: *Chemistry Atoms First* [OpenStax](#)

COVID-19 Statement:

Although COVID-19 has reached an endemic phase it is still important to remain vigilant as we face a recent rise in positive cases. As socially responsible members of this community, everyone is expected to engage in daily health self-monitoring, to stay home (residence hall or off-campus housing) from on-campus class, work, or activities if they begin experiencing any COVID-related symptoms.

When experiencing any COVID-related symptoms, students are expected to contact Health Services by completing the QI form in the [Patient Portal](#) and respond to the nurse who will contact them with instructions. COVID positive residential students are required to follow their QI plan for 5 days of isolation off campus so be prepared with a back-up plan as well.

By acknowledgement, you agree to Winthrop's expectations of you regarding health monitoring and reporting.

Course Goals:

- In this lab recitation course, we will review the necessary information needed to successfully complete each of the experiments scheduled in CHEM 204. Many of the laboratory skills learned in CHEM 204 will be used in upper level chemistry labs.
- Develop problem-solving and critical thinking skills.
- Demonstrate an understanding of the fundamental principles presented in each of the laboratory experiments conducted in CHEM 204.

Course Outline:

- Physical properties, analytical balances, typical glassware, volumetric glassware, data analysis, graphing using Excel
- Qualitative solubility, stoichiometry, volumetric glassware, burets, pH meters, writing chemical equations
- Calorimetry, heats of reactions, endothermic/exothermic reactions, writing chemical equations, Excel
- Introduction to organic chemistry, identifying an unknown, functional group testing, infrared spectroscopy, NMR
- Visible spectroscopy, Beer's Law, calibration curves, graphing with Excel

Exams and Grading:

1. There will be two exams each worth 60 points. **No make-up exams will be given.** Tentative exam dates are noted below.
2. Your final grade in lab recitation will be factored into your final CHEM 204 grade. Recitation will account for about 20% of your final lab grade. See the CHEM 204 GRADING webpage for letter grade assignments.

3. There will not be a final exam during final exam time in recitation. There will be a cumulative exam (Exam 2) for recitation during the last week of classes.
4. You have one week from the time a graded assignment is returned to question its grading. After a week, I will not change any grade.
5. You should carefully read the Winthrop University Student Conduct Code printed in the Winthrop University Student Handbook. As noted in the Student Conduct Code: "Responsibility for good conduct rests with students as adult individuals." This policy on student academic misconduct is outlined in the **Student Conduct Code Academic Misconduct Policy** (<https://www.winthrop.edu/uploadedFiles/studentconduct/StudentConductCode.pdf>) in the online *Student Handbook* <https://www.winthrop.edu/studentconduct/winthrop-university-student-handbook.aspx>

Total Possible Points

Recitation Exams	120 pts (20%)
<u>Laboratory grade</u>	<u>480 pts (80%)</u>
Total points Chem 204	600 pts

Exams: You will need a calculator for exams. Cell phones and pagers are strictly prohibited during exams. You cannot use a cell phone as a calculator during exams.

Course Withdraw: Tuesday, March 19th is the last day to withdraw from a full fall semester course with an automatic N grade issued. Students may not withdraw from a course after this date without documented extenuating circumstances as determined by the University.

Communication: If you have any questions, please stop by and see me during office hours. If these hours are not convenient, see me in class or e-mail me to set up an appointment.

Attendance: You are expected to attend all class meetings. You are responsible for all announcements made in class. Absence or lateness does not excuse you from this responsibility.

Homework: You are expected to complete each reading assignment and all assigned homework problems. End of chapter homework problems from the textbook will not be collected or graded, but will be good practice for preparing for exams.

E-mail: It is important to check your e-mail regularly. If you registered for the course late, you will need to manually subscribe to the listserv. If you drop the course, you will need to unsubscribe to the list or you will continue to receive all e-mails I send. You can find directions at <http://www.winthrop.edu/acc/classlist.htm>

Students with Disabilities: Winthrop University is committed to providing access to education. If you have a condition which may adversely impact your ability to access academics and/or campus life, and you require specific accommodations to complete this course, contact the Office of Disability Services (ODS) at 803-323-3290, or, accessibility@winthrop.edu. Please inform me as early as possible, once you have your official notice of accommodations from the Office of Disability Services.

Syllabus: Mistakes to this syllabus will be corrected as necessary.

This is a tentative schedule and will be modified as needed.

Date	Exams	Prelab Topics	Reading Assignments and Homework Problems (End of Chapter Exercises)	
1/9, 1/16		Quantitative and Qualitative Analysis, Density, Units of Measurement, Making Measurements: Precision, Accuracy, Experimental Error, Standard Deviation, Significant Figures	<p>Chemistry: An Atoms-Focused Approach: ~Read Section 1-4, Density; Sections 1-8 and 1-9 ~Complete the following problems:</p> <ul style="list-style-type: none"> • End of Chapter Problems 63, 64, 65 <p>A sample of an unknown metal was placed in a graduated cylinder containing water. The mass of the sample was 23.5 g and the water level rose from 47.5 ml to 52.2 ml. Calculate the density of this unknown metal.</p>	<p>Chemistry Atoms First OpenStax ~Read Chapter 1 ~Complete the following problems:</p> <ul style="list-style-type: none"> • End of Chapter Problems 55, 81, 87, 89 • A sample of an unknown metal was placed in a graduated cylinder containing water. The mass of the sample was 23.5 g and the water level rose from 47.5 ml to 52.2 ml. Calculate the density of this unknown metal.
1/23		Ions, Ionic compounds, Precipitation reactions, Solubility rules, Qualitative analysis	<p>Chemistry: An Atoms-Focused Approach: ~ Ions, Ionic Compounds: Read pages 43, 50-53, 139-142 ~Reactions: Read Section 8.5 including Sample Exercise 8.6. Complete Practice Exercise on page 328 and End of Chapter Problems 8.67, 8.68</p>	<p>Chemistry Atoms First OpenStax ~ Ions, Ionic Compounds: Read Section 3.7 ~Reactions: Read Section 7.2 Precipitation Reactions including Example 7.3. Exercise 11.</p>

1/30, 2/6		Writing chemical equations for precipitation reactions (complete balanced equations, complete ionic equations, and net ionic equations), Acid/Base reactions, Molarity, Dilutions, Using burets, Using pH meters Stoichiometry	<i>Chemistry: An Atoms-Focused Approach:</i> ~ Molarity: Read Section 8.1 (312-315); Complete End of Chapter Problems 8.11a; 8.14a,b;15a,b ~Acids and Bases: Read Section 8.4 (320-326) including Sample Exercise on page 324; Complete End of Chapter Problems 8.51b, 8.53b ~Reaction Stoichiometry: Read Section 8.5, (326-330); Read Section 8.5 including Sample Exercise 8.7 and 8.8, complete Practice Exercises on page 330 and 331 and End of Chapter Problems 8.68, 8.69, 9.72	<i>Chemistry Atoms First OpenStax</i> ~ Molarity: Read Section 6.3; Complete End of Chapter Problem 22 ~Acids and Bases: Read Section 7.2 ~Reaction Stoichiometry: Read Section 7.3; including Examples 7.10 and 7.11, 7.13 and End of Chapter Exercises 7.50, 7.63, 7.64
2/13, 2/20, 2/28		Calorimetry, Heats of Reactions Precipitation Reactions, Acid/Base Reactions, Oxidation-Reduction Reactions	<i>Chemistry: An Atoms-Focused Approach:</i> ~ Calorimetry: Read Section 9.5, (382-384) including sample Exercise (384); Complete Practice Exercise on page 384 and End of Chapter Problems 9.65, 9.66 ~Precipitation Reactions: Read Section 8.5 ~ Acids and Bases: Read Section 8.4 ~Oxidation-Reduction Reactions: Read Section 8.6, (332-338); Complete End of Chapter Problems 8.83, 8.89	<i>Chemistry Atoms First OpenStax</i> ~ Calorimetry- Read Section 9.2 including examples 9.5 and 9.6; Complete End of Chapter Problem 24, 25, 26.
3/5	Exam 1			
3/12	No Class	Spring Break		
3/19		Organic Nomenclature and Functional Groups NMR and IR		
3/26, 4/2		Analysis of Cola	<i>Chemistry: An Atoms-Focused Approach:</i> ~Determining Concentration (343-344)	
4/9		Review		
4/16	Exam 2			