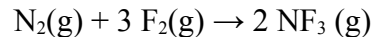


CHEM105 Announced Quiz 6

1. Use bond enthalpies to estimate the reaction enthalpy for:



2. A 10.0 g piece of copper at 100°C was placed into a container having negligible heat capacity containing 60.0 mL of water at 25.0°C. Determine the final temperature of the water. The heat capacity of copper is 0.38 J / g - °C.
3. The enthalpy of combustion for propane, C₃H₈, is 2220 kJ / mole.
- Write a balanced chemical equation for the combustion of propane.
 - Calculate the amount of heat released from the combustion of 22.0 grams of propane.