

Quiz5Key

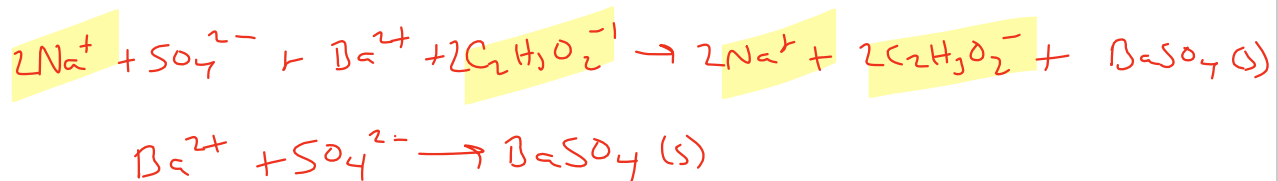
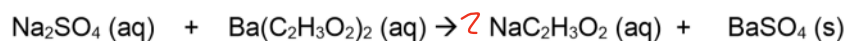
Thursday, October 13, 2016 2:29 PM

Name Key

1. How many moles of sodium are found in 12 L of 6.3 M Na_2SO_4 ?

$$\frac{12 \text{ L} \times 6.3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \times \frac{2 \text{ mol Na}}{1 \text{ mol Na}_2\text{SO}_4} = 151.2 \text{ mol Na}^+$$

2. Write a net ionic equation for the following reaction.



Identify the spectator ions.



3. Explain why NaCl is soluble in water. The ion-dipole interactions between the sodium cation / chloride anion and water are stronger than the individual interactions between Na-Cl and $\text{H}_2\text{O-H}_2\text{O}$.