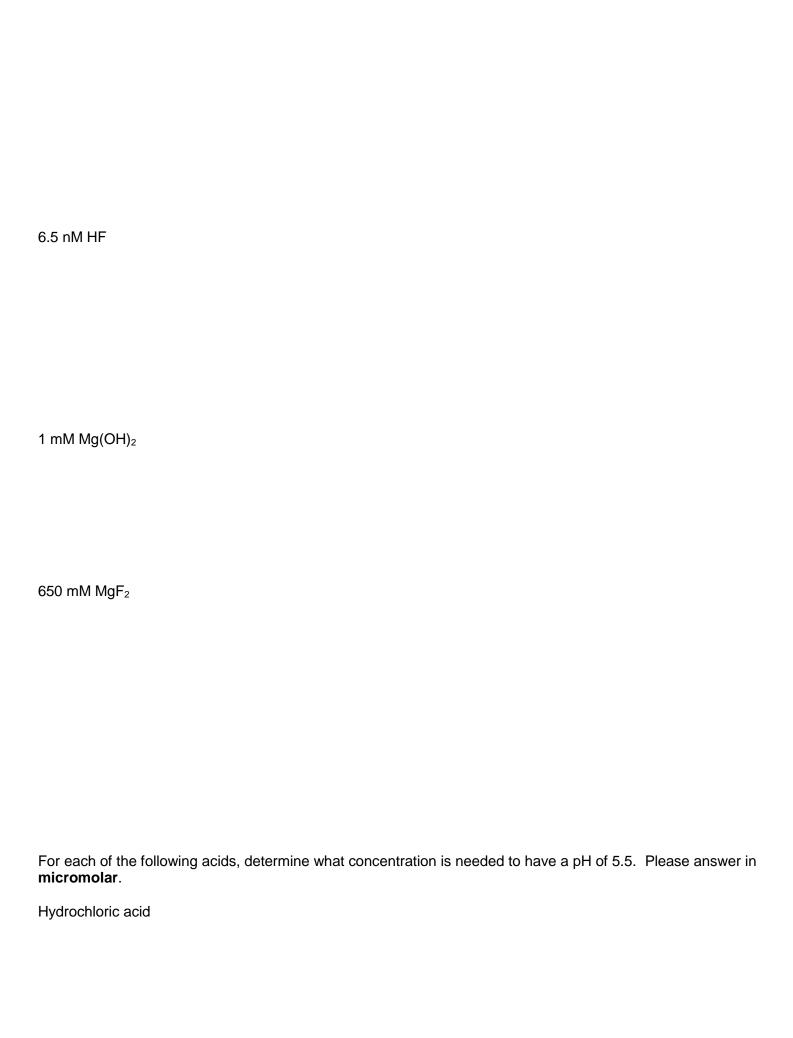
Acids and Bases

For each of the	following salts, p	redict if a 100) mM solution wou	ld be acidic, basic, o	or neutral.	
	NaCl	KNO ₂	sodium acetate	sodium hy	drogen sulfate	
			or base? How about so it is NOT a base. N		+ Cl ⁻ . Na ⁺ is not a protor	n donor
For each pair, id	dentify which will	be a more ac	idic solution:			
10 mM HCl or 1	0 mM HF	10 mM F	I₂SO₄ or 10 mM H	CI 10 mM H	NO ₂ or 20 μM HNO ₂	
					concentration, acid streng nd HF is a weak acid. HC	
Order the follow	ving solutions by i	ncreasing ac	idity (lowest pH go	pes last). A table of	pKa is attached.	
100	mM HF,	100 mM HC	CIO, 100 mN	И HSO ₄ -1,	100 mM NH ₄ +	
			a 100 mM NaHC0 a stronger acid or bas	O_3 solution, will it be e ?	acidic or basic?	
Hypochlorous a and Kb of the co		pKa of 7.53.	What is the Ka?	What is the conjuga	ite base? What are tl	he pKb
Calculate the pl	H of each of the fo	ollowing solut	tions:			
650 μM HCI		J				

 $650~\mu\text{M}$ HF



Ammonium chloride